



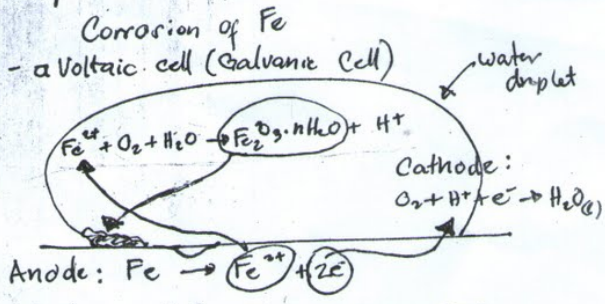
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Experiment 7



What's needed for corrosion of Fe?
 H_2O
 O_2 } must be present together

H^+ - catalyst; depleted then regenerated
 salts - enhances conductivity of the water thus promotes corrosion

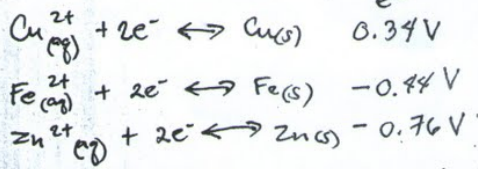
Acidic solutions (x's H^+)
 - will give visual indication of corrosion after 1 day
 - will give blue ppt with $KFe(CN)_6$

Neutral solutions (ions present)
 - (+) indication (visual) in 1 day
 - will give blue ppt with $KFe(CN)_6$

Basic solutions (x's OH^-)
 - depletes H^+ in the solution thus preventing $O_2 + H^+ + e^- \rightarrow H_2O$ from happening
 - no indication of corrosion in 1 day
 - no reaction with $KFe(CN)_6$

* $KFe(CN)_6$ turning blue indicates Fe^{2+} (corrosion)

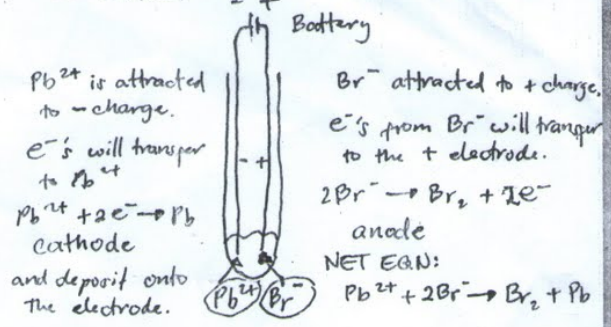
- Nail Only - corrosion on both ends due to stress/scratches (increased surface area)
- Bent Nail - corrosion on both ends and on the bend due to stress marks
- Nail w/ Zn - Zinc corroded to Zn^{2+} . Fe is untouched. Zn is more likely to lose e^- 's because of lower E°
- Nail w/ Cu - Cu is less active than Fe. Fe is more likely to lose e^- 's because it has lower E° than Cu.



The more positive E° becomes the cathode reaction.

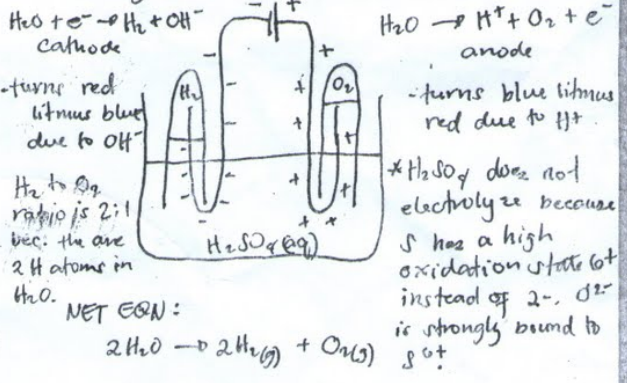
Experiment 8 Electrolysis of $PbBr_2$

- Salts are nonconductive unless in a solution or molten

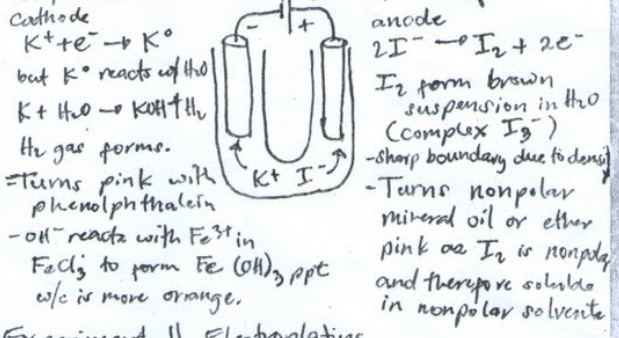


Experiment 9. Electrolysis of H_2SO_4

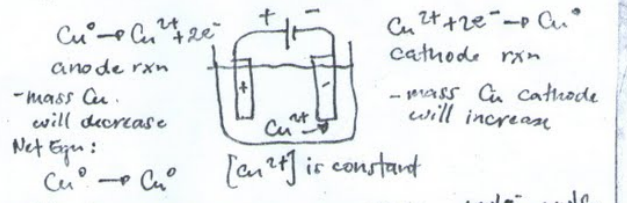
- is actually electrolysis of H_2O with H_2SO_4 as catalyst. (H_2SO_4 increases as H_2O is consumed.)



Experiment 10. Electrolysis of KI(aq)



Experiment 11 Electroplating



T(s)	A	Ave I	ΔT	Q	107.10 x (1 mole e- / 96500 C) x (1 mol Cu / 2 mole e-)
0	0.25	0	0-180	36	0.5875g Cu
180	0.17	0.2	180-360	27.7	0.035g
360	0.14	0.15	360-540	23.7	
540	0.12	0.13	540-720	19.8	
720	0.1	0.1	720-900	107.1	

* The anode should decrease by 0.035g and the...

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